

Name: _____

Period: _____

Seat#: _____

Show work and include ALL units.

Use single dimensional analysis line method setups for conversions!

1) 14.8g of C_3H_8 and 3.44 g of O_2 react in the following reaction:



- Determine the limiting reagent and the excess reagent
- Determine the number of grams of H_2O produced
- Determine the number of grams of excess reagent left

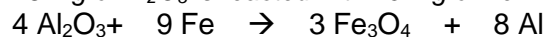
2) 10.0 g of $Al_2(SO_3)_3$ is reacted with 10.0 g of $NaOH$ in the following reaction:



- Determine the limiting reagent and the excess reagent
- Determine the number of grams of Na_2SO_3 produced
- Determine the number of grams of excess reagent left

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3) 25.4 g of Al_2O_3 is reacted with 10.2 g of Fe:



- Determine the limiting reagent and the excess reagent
- Determine the number of moles of Fe_3O_4 produced
- Determine the number of grams of excess reagent left

- 4) When copper (II) chloride reacts with sodium nitrate, copper (II) nitrate and sodium chloride are formed.
- Write the balanced equation for the reaction given above.
 - If 15 g of copper (II) chloride react with 20 g of sodium nitrate what is the limiting reagent for the reaction?
 - How much sodium chloride can be formed in grams?
 - How many grams of copper (II) nitrate is formed?
 - How many grams of the excess reagent are left over in this reaction?
 - If 11.3 g of sodium chloride was actually formed in the reaction, what is the percent yield of this reaction?

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- 5) 1000 grams of sodium chloride is combined with 2000 grams of barium phosphate
- How many grams of each product are made?
 - How many grams of the excess reagent are left over in this reaction?

- 6) Iodine gas reacts with Calcium fluoride.
- How many grams of your calcium containing compound do you have when this reaction is over if you started with 140g of Iodine gas, and 3.75 moles of calcium fluoride?
 - How many grams of the excess reagent do you have left?